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January 2010

FIGURE 1. Salts, acids, and bases are electrolytes. They dissolve in water to form ions. ... In practice, the cell constant is measured against a solution of known conductivity. The cell constant is the ratio of the known conductivity (µS/cm) to ...

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GLP (good laboratory practice) features allow the user to store and retrieve data regarding the status of the system. And for those on a tight budget, some conductivity bench meters will even include a pH meter with two separate electrodes. These combination meters share the same display. Process Conductivity/TDS Controllers

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Conductivity (or specific conductance) of an electrolyte solution is a measure of its ability to conduct electricity. The SI unit of conductivity is Siemens per meter (S/m). Conductivity measurements are used routinely in many industrial and environmental applications as a fast, inexpensive and reliable way of measuring the ionic content in a solution. For example, the measurement of product ...

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sufficient conductivity. As described in 2.2 the reference potential is constant as long as the internal electrolyte is constant. The reference potential also varies slightly with temperature. The tube containing all the reference elements and solutions/gel is in contact with the sample to measure through a junction (diaphragm).

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Therefore, conductivity is defined as the reciprocal of resistance of a solution between two electrodes and it can be expressed as: G = I / R. The basic unit of conductivity is the siemen (S), formerly called the mho. Since the cell geometry determines the conductivity value it can be used to measure, standardized

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Welcome to the Conductivity Theory Page. Conductivity is the ability of a solution to conduct electric current. The principle by which instruments measures conductivity is simple - two plates (cells) are placed in the sample, a potential is applied across the plates and the current is measured.
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